

Molecular Spectroscopy: Hilary Term 2003
Questions

- (1)(a) Which of the following molecules would you expect to show a microwave (pure rotational) spectrum: H_2 , HCl , CH_4 , CH_3Cl , CH_2Cl_2 , H_2O , N_2 , H_3 , benzene, toluene, CO_2 ?
- (b) Suggest a structure for XY_3 so that it (a) gives a rotational spectrum, (b) does not give a rotational spectrum (Give reasons).
- (2) The molecule H^{35}Cl exhibits rotational absorption lines in the far infrared at the following wavenumbers (cm^{-1}): 83.32, 104.13, 124.73, 145.37, 165.89, 186.23, 206.60, and 226.86. (Note that there may be other lines in the microwave region too).
- (a) Identify the transitions and use a graphical procedure to determine the rotational and centrifugal distortion constants. Calculate the bond length of HCl .
- (b) Predict the rotational constant for DCl .
- (c) Determine the most populated rotational level in HCl at 300K.
- (3) Given that the CO bond length in the molecule OCS is 0.1165 nm and the CS bond length is 0.1558 nm, predict the frequencies (units Hz) of the $J = 1 \leftarrow 0$ and $2 \leftarrow 1$ transitions in the rotational spectrum of OCS .
- (4) Calculate the energies of the $J = 2$ levels of CF_3I , ignoring centrifugal distortion, given $A_0 = 0.191 \text{ cm}^{-1}$ and $B_0 = 0.0507 \text{ cm}^{-1}$. What would be the magnitudes of the Stark splittings of these levels in a static electric field of 10 Vm^{-1} if the dipole moment of CF_3I is $3.34 \times 10^{-30} \text{ Cm}$ (1 Debye). Draw an energy level diagram illustrating the splittings.
- (5)(a) The energy levels of a symmetric top molecule including centrifugal distortion are given (in wavenumber units) by the expression
- $$F(J,K) = BJ(J+1) + (A-B)K^2 - D_J J^2(J+1)^2 - D_{JK} J(J+1)K^2 - D_K K^4$$
- where D_J , D_{JK} and D_K are centrifugal distortion constants. Derive an expression for the wavenumbers of the allowed transitions in the rotational $\nu = 0$ spectrum ($\Delta J = \pm 1$, $\Delta K = 0$)
- (b) In the rotational spectrum of CH_3F , lines were observed at the frequencies 51071.8, 102142.6, 102140.8, 153210.3, 153207.6, 153199.6 MHz. Assign quantum numbers to the transitions and determine B , D_J , and D_{JK} for CH_3F .
- (6) What is the number of vibrational modes for the following molecules?
 NH_3 , HCN , SO_2 , C_2H_2 , C_{10}H_8 , C_{60} .
- Sketch the normal modes for HCN and SO_2
- (7) The following lines are observed in the fundamental (infrared) vibration-rotation band of H^{35}Cl : 2998.05, 2981.05, 2963.35, 2944.99, 2925.92, 2906.25, 2865.14, 2843.63, 2821.59, 2799.00, 2775.77 cm^{-1} .
- (a) Assign the rotational quantum numbers for the $\nu = 1$ and $\nu = 0$ levels associated with each transition.
- (b) Graphically or otherwise, using the method of combination differences, determine the rotational constants for both levels (B_1 and B_0); ignore centrifugal distortion.
- (c) Hence determine the equilibrium internuclear distance (r_e).
- 8 (a) Derive an expression for the separation between adjacent lines in the P-branch of the fundamental vibration-rotation band of a diatomic molecule in terms of the rotational constants B_1

and B_0 (ignore centrifugal distortion).

(b) Lines are observed in the P branch of the fundamental band of carbon monoxide with the following wavenumbers / cm^{-1} .

P(9)	P(10)	P(14)	P(15)
2107.425	2103.271	2086.323	2082.003

Determine the rotational constants B_1 and B_0 and comment on their relative values.

(9) Adjacent vibration-rotation lines near the centre of the ν_3 stretching fundamental band in the infrared spectrum of CO_2 occur at the following wavenumbers:

2351.64, 2350.08, 2347.74, 2346.18 cm^{-1} .

Calculate the band origin of this fundamental vibration and the C–O internuclear distance.

(10)(a) A band in the infrared spectrum of ethyne shows P, Q and R branches. Draw an energy level diagram of the rotational levels in the two vibrational states and use it to explain why you might expect the frequency differences $\Delta_1 = R(J) - Q(J)$ and $\Delta_2 = Q(J+1) - P(J+1)$ to be the same.

(b) The measured values of Δ ($= \Delta_1 = \Delta_2$) are given for some J values below.

J	Δ / cm^{-1}
5	14.144
10	25.919
15	37.676
20	49.404
25	61.096

Extract as many parameters as possible from these data. What information do they provide about the structure of the molecule?

(11) Determine the symmetries of the vibrational modes for a homonuclear triatomic molecule X_3 with an equilateral triangular structure (D_{3h} symmetry) and hence determine which vibrations are active in (a) the infrared and (b) the Raman spectrum.

Ozone, O_3 exhibits fundamental vibrational bands in the infrared spectrum at 710 cm^{-1} , 1043 cm^{-1} and 1103 cm^{-1} . Is this observation consistent with the hypothesis that ozone has an equilateral triangular structure?

Consider one other possible structure and determine the infrared and Raman rules for it. Is this structure consistent with the experimental evidence given above?

In the infrared spectrum of BF_3 , the A_2'' fundamental ν_2 and its second overtone $3\nu_2$ are seen in absorption from the ground state, but not the first overtone $2\nu_2$. Why is this?

(12) (a) The rotational Raman spectrum of $^{35}\text{Cl}_2$ shows a series of lines separated by 0.9752 cm^{-1} in both Stokes and anti-Stokes branches. Determine the bond length of Cl_2 .

(b) Given that the CO and CS bond lengths in CO_2 and CS_2 are 0.1162 nm and 0.1555 nm respectively, explain why the line spacing in the rotational Raman spectrum of OCS is very close to that for CS_2 .

(13) The rotational Raman spectrum of H_2 was recorded at 350K . The displacements of the first 5 lines of the Stokes branch from the exciting line are listed below, together with their relative intensities.

Displacement / cm^{-1}	Intensity
364.8	1.0
608.0	5.46
851.2	1.12
1094.4	1.05
1337.6	0.06

(a) Explain *quantitatively* the magnitude of both the displacements and intensities.

(b) Predict what differences in displacements and intensities you would expect for D_2 .

(14) The table below lists the vibrational wavenumbers and infrared/Raman activities for cyanogen (C_2N_2). What does this information tell us about the structure of cyanogen? Suggest an assignment for the bands.

226 cm^{-1}	Infrared	PQR branches
506 cm^{-1}	Raman	
848 cm^{-1}	Raman	
2149 cm^{-1}	Infrared	PR branches only
2322 cm^{-1}	Raman	

The rotational Raman spectrum of the molecule shows a series of anti-Stokes lines separated by 0.63 cm^{-1} . If the CN bond length is 0.1165 nm , determine the C–C bond length.